De-pyritization of Tailings: Experiences from Three Copper Mines in Scandinavia



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Definition and legal aspects

Definition of de-pyritization according to the EU:

• "Separation of pyrite from the tailings and separate discharge of the pyrite"

EU- legal requirements:

The management of extractive waste should be based, inter alia, on the concept of BAT

BAT on ARD management:

Characterization of extractive waste and determination of the acid-forming potential.

If ARD-forming potential exists:

1. Prevent the generation of ARD (de-pyritization is listed as a prevention option),

2.<u>Control</u> ARD

3. Apply treatment options.

Often a combination is used.

The de-pyritization technique is well known. The pyritic product has a high reactivity and therefore carefully designed measures for deposition are required.



General objectives

General objectives of performing de-pyritization are often:

- Avoid/lower ARD potential in the de-pyritized tailings
- Concentrate potentially economically recoverable substances
- Legal requirements
- Permit requirements

Potential benefits:

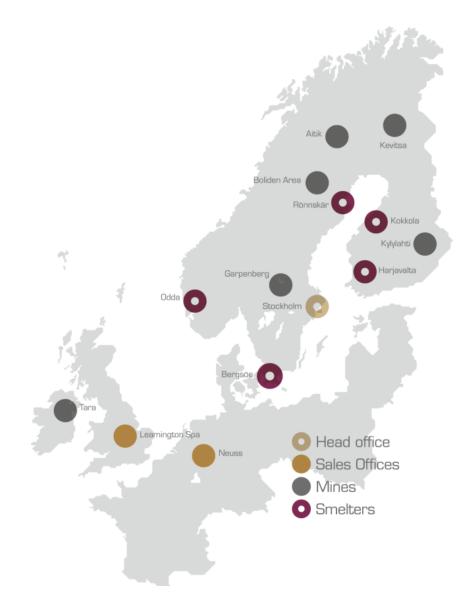
- Minimize long-term risks
- Recover more from the ore
- Lower closure costs
- Alternative closure options
- Permitting



Boliden sites

The following Boliden sites perform de-pyritisation:

- Kevitsa, Finland
- Kylylahti, Finland
- Aitik, Sweden





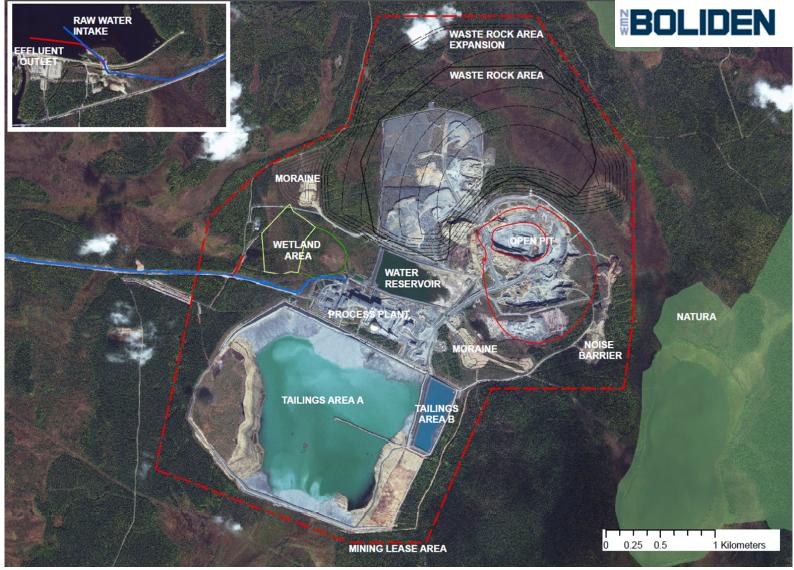
KEVITSA

Finnish nickel-copper-gold-Platinum Group Metals (PGM) mine

BOLIDEN

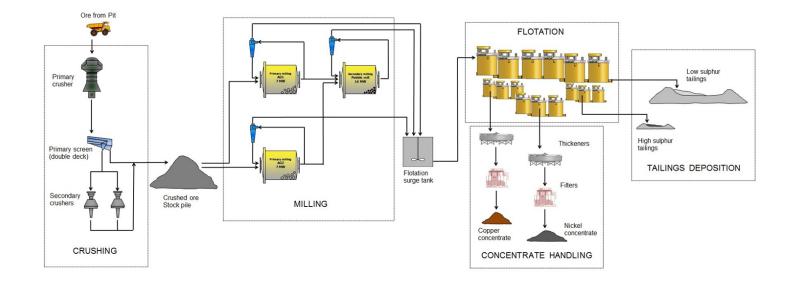
- Milled tonnage: 7 -9 Mtonnes
- Waste-rock to ore ratio: 5:1
- Number of employees 380 and 200 contractors

Kevitsa





Kevitsa performance



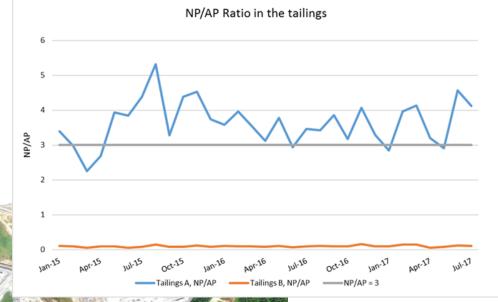


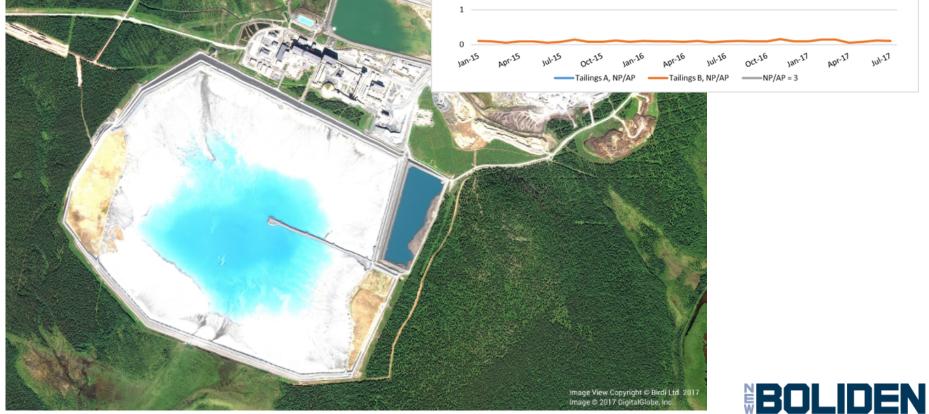
Kevitsa performance



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Kevitsa performance





KYLYLAHTI – Copper mine

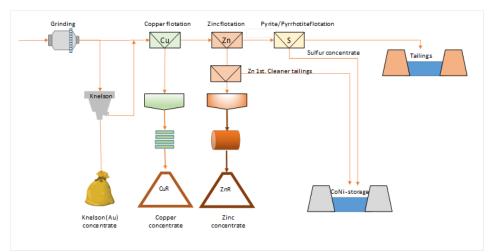
- Finnish copper, gold, zinc and silver underground mine
- Exploration rights in the surrounding Outokumpu field
- Milled tonnage: 800 Ktonnes
- Number of employees 110



Kylylahti

- General
 - Underground copper mine in eastern Finland. Luikonlahti concentrator.
 - Annual production 800 kton.
 - Cu, Zn and Au concentrates.
 - 470 kton tailings, 35 kton CoNi tails and 230 kton S-tails
- Objectives
 - To reduce S-concentration to acceptable level
 - To concentrate Co, Ni for future process
- Permit requirements
 - Tailings should not be significantly acid forming (not non-acid forming)
 - CoNi+S tailings to lined facility
- Closure
 - Dry cover
 - Closure Plan under revision

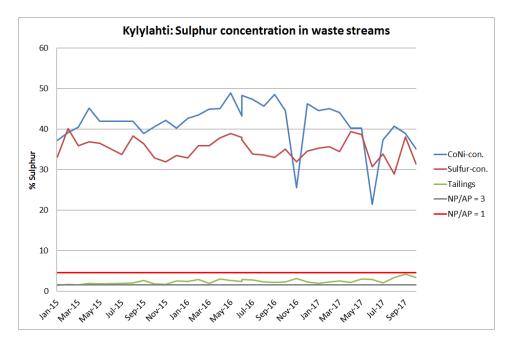
Schematic Process layout





Kylylahti performance







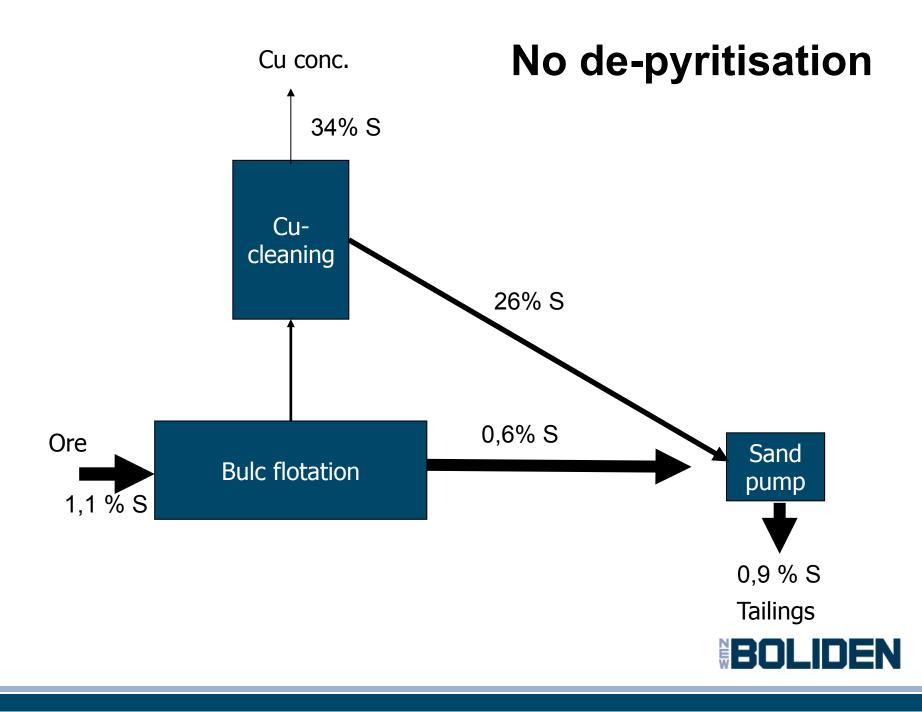
AITIK –open-pit copper mine

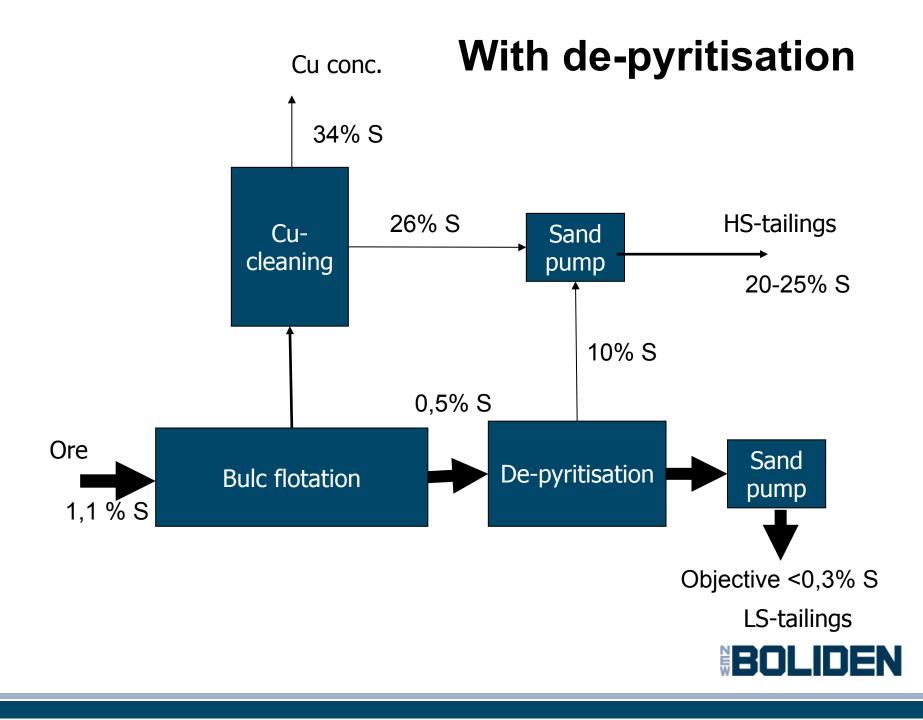
- Swedish copper, gold, silver open pit mine
- Large-scale operations
- Milled tonnage: 39,000 Ktonnes
- Waste-rock to ore ratio: 1:1
- Number of employees 700



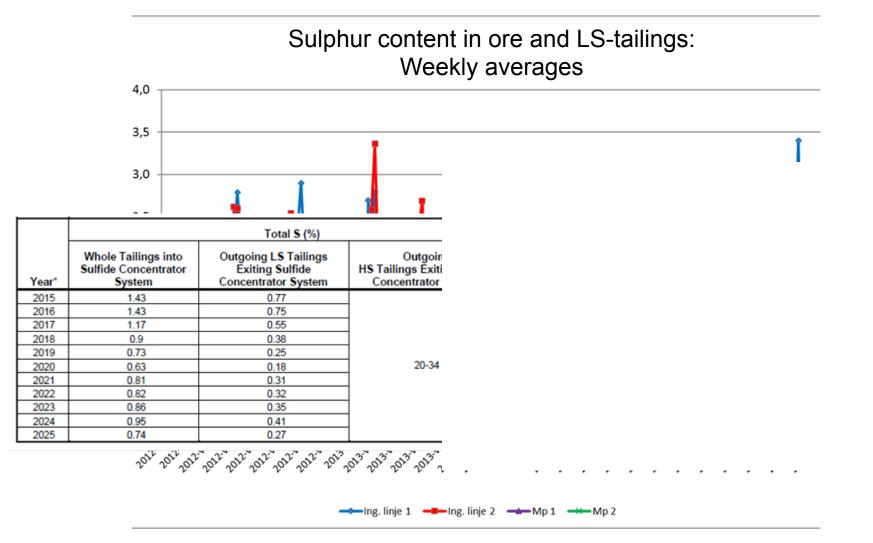








Aitik performance

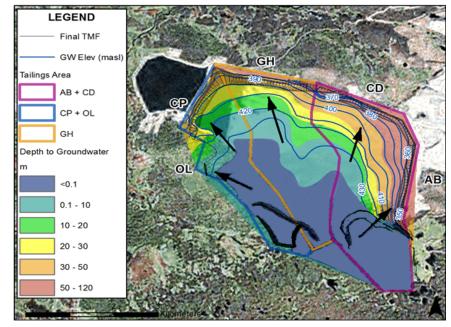




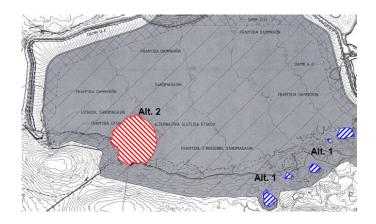
Aitik – site evolution

Changing conditions over time:

- Dam construction methods
- Hydro-geology
- Potential thiosalt generation in HS-pond
- Permitting process



From: Hatch, 2015





BOLIDEN

Discussion

Overall balance:

when and why de-pyritization may be worth while and for which reasons?



Issues

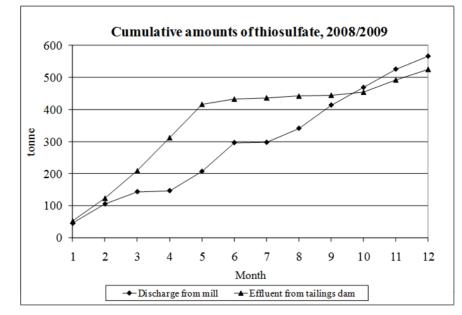
- Varying performance due to incoming Sgrades and production rates
- Objectives how to evaluate them

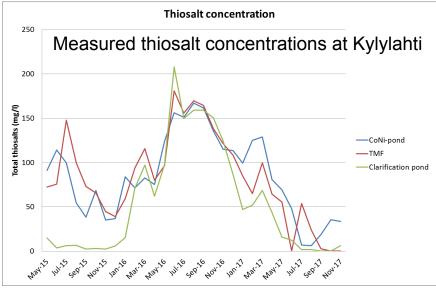
Segregation

- Potential thiosalt generation in high sulphur tailings at high pH
- Closure (both for LS and HS)

Evolving site

- Evolving environmental requirements non-ARD not enough
- Increased foot-print
- Re-processing opportunities/problemsCosts



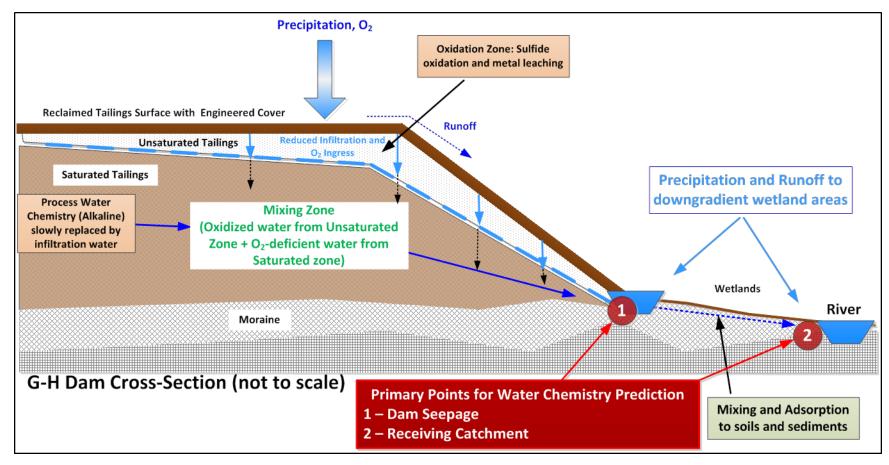








Aitik – site evolution



From: Hatch, 2015



Thiosalts Generation

- Disproportionation of sulphur by hydroxide ions; dominating in the tailings pond?

By the action of alkalinity on elemental sulphur (eqs 1-2).

$$4 S^{0} + 4 OH^{-} \rightarrow S_{2}O_{3}^{2^{-}} + 2 HS^{-} + H_{2}O$$
(1)

 $4 \text{ S}^0 + 5 \text{ OH}^- \rightarrow \text{SO}_4^{2-} + 3 \text{ HS}^- + 4 \text{ H}_2\text{O}$

By the action of alkalinity on pyrite (eqs 3-4).

- $4 \text{ FeS}_2 + 4 \text{ OH}^- \rightarrow \text{S}_2\text{O}_3^{2-} + 2 \text{ HS}^- + \text{H}_2\text{O} + 4 \text{ FeS}$ (3)
- $4 \text{ FeS}_2 + 5 \text{ OH}^- \rightarrow \text{SO}_4^{2-} + 3 \text{ HS}^- + 4 \text{ H}_2\text{O} + 4 \text{ FeS}$ (4)



(2)